

CHEMISTRY

FORMULA NAMING AND WRITING (IONIC & MOLECULAR)

A. Directions: Use your table of common ions and the periodic table to classify the following compounds as either molecular (M) or ionic (I) and then name the compounds. Remember to use the Greek prefix system when it is appropriate.

	I or M	Name
1. CaBr ₂	_____	_____
2. CuO	_____	_____
3. N ₂ O ₅	_____	_____
4. Li ₃ PO ₄	_____	_____
5. FeSO ₄	_____	_____
6. CCl ₄	_____	_____
7. AgNO ₃	_____	_____
8. SO ₃	_____	_____
9. SnF ₄	_____	_____
10. Al(OH) ₃	_____	_____

B. Directions: Use your table of common ions and the periodic table to classify the following compounds as either molecular (M) or ionic (I) and then write the formulas for the compounds.

	I or M	Formula
11. Strontium Bromide	_____	_____
12. Lithium Carbonate	_____	_____
13. Nitrogen Trifluoride	_____	_____
14. Ferric Phosphate	_____	_____
15. Aluminum Oxide	_____	_____
16. Lead II Hydroxide	_____	_____
17. Tetraphosphorus decoxide	_____	_____
18. Barium Chlorate	_____	_____
19. Chromium III Carbonate	_____	_____
20. Potassium Permanganate	_____	_____