1.   Identify the oxidizing agent in
      $Al + H^+ \longrightarrow Al^{3+} + H_2$.

                                          [1] _____

2.   Identify the reducing agent in
      $H_2SO_4 + MnO_2 \longrightarrow SO_4^{2-} + Mn^{2+}$.   [2] _____

3.   Find the oxidation number of the underlined atom in
      a) $H_2\underline{S}$.
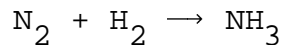   b) $\underline{I}O_4^{-}$.

                                          [3] _____

4.   Is the following an oxidation-reduction reaction? If yes, name the
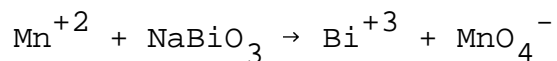   element reduced.
      $Br_2 + SO_3^{2-} \longrightarrow Br^- + SO_4^{2-}$
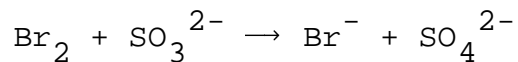
                                          [4] _____

5. Is the following an oxidation-reduction reaction? If yes, name the element oxidized.
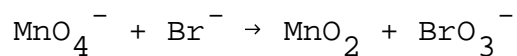
N_2 + H_2 $\longrightarrow$ NH_3

[5] _____

6. Balance the following oxidation-reduction reaction in acidic solution. How many H_2O molecules are needed?

Mn^{+2} + $NaBiO_3$ → Bi^{+3} + MnO_4^-

[6] _____

7. Balance the following oxidation-reduction reaction in basic solution. How many H_2O molecules are there?

MnO_4^- + Br^- → MnO_2 + BrO_3^-

[7] _____