

REACTION CATEGORY	DECOMPOSITION
REACTION DESCRIPTION	During decomposition, one compound splits apart into two or more substances. These substances can be elements or simpler compounds.
REACTION FORMAT	$AB \rightarrow A + B$
REACTION GUIDELINES	<ol style="list-style-type: none"> 1. Binary compounds breakdown into their elements. 2. Carbonates break down into an oxide and carbon dioxide 3. Chlorates break down to a binary salt and oxygen. 4. Bases break down to the oxide of the metal and water. 5. Acids break down to the oxide of the nonmetal plus water.
REACTION GUIDELINE EXAMPLES	<ol style="list-style-type: none"> 1. $2 \text{NaCl} \rightarrow 2 \text{Na} + \text{Cl}_2$ 2. $\text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{O} + \text{CO}_2$ 3. $\text{Ba}(\text{ClO}_3)_2 \rightarrow \text{BaCl}_2 + \text{O}_2$ 4. $\text{Ca}(\text{OH})_2 \rightarrow \text{CaO} + \text{H}_2\text{O}$ 5. $2 \text{H}_3\text{PO}_4 \rightarrow \text{P}_2\text{O}_5 + 3 \text{H}_2\text{O}$

Decomposition Reaction Practice

