

Identify each of the following chemical equations as either:

[A] direct combination or synthesis

[B] decomposition or analysis

[C] single replacement

[D] double replacement

[E] complete combustion

- ____1] CO_2 -----> $\text{C} + \text{O}_2$
- ____2] $\text{NaCl} + \text{AgNO}_3$ -----> $\text{NaNO}_3 + \text{AgCl}$
- ____3] $\text{S} + \text{Cl}_2$ -----> SCl_2
- ____4] $\text{BaCl}_2 + 2 \text{NaOH}$ -----> $2 \text{NaCl} + \text{Ba(OH)}_2$
- ____5] $\text{Zn} + \text{CuSO}_4$ -----> $\text{ZnSO}_4 + \text{Cu}$
- ____6] $\text{C}_4\text{H}_{10} + 13 \text{O}_2$ -----> $8 \text{CO}_2 + 10 \text{H}_2\text{O}$
- ____7] $\text{Pb(NO}_3)_2 + \text{Mg}$ -----> $\text{Pb} + \text{Mg(NO}_3)_2$
- ____8] $\text{Mg} + 2 \text{HCl}$ -----> $\text{MgCl}_2 + \text{H}_2$
- ____9] H_2SO_4 -----> $\text{H}_2 + \text{S} + 2\text{O}_2$
- ____10] $2 \text{O}_2 + \text{N}_2$ -----> N_2O_4
- ____11] $3 \text{CaBr}_2 + 2 \text{Na}_3\text{P}$ -----> $\text{Ca}_3\text{P}_2 + 6 \text{NaBr}$
- ____12] $2 \text{KI} + \text{Br}_2$ -----> $2 \text{KBr} + \text{I}_2$
- ____13] $\text{C}_6\text{H}_{12}\text{O}_6$ -----> $6 \text{C} + 6 \text{H}_2\text{O}$
- ____14] 2NaF -----> $2 \text{Na} + \text{F}_2$
- ____15] $\text{Si} + \text{O}_2$ -----> SiO_2
- ____16] $2 \text{NaI} + \text{Pb(NO}_3)_2$ -----> $2 \text{NaNO}_3 + \text{PbI}_2$
- ____17] $\text{NaI} + \text{Cs}$ -----> $\text{CsI} + \text{Na}$
- ____18] $\text{H}_2 + \text{CO} + \text{O}_2$ -----> H_2CO_3
- ____19] Li_3PO_4 -----> $3 \text{Li} + \text{P} + 2 \text{O}_2$
- ____20] $\text{CS}_2 + 2 \text{F}_2$ -----> $\text{CF}_4 + 2\text{S}$